James Clemens High School
11306 County Line Road
Madison, AL 35756

Phone: 256-216-5313
Extension: 95237
Email: ahorne@madisoncity.k12.al.us

Course Syllabus

Chemistry I Fall 2022

Instructor: Ashley Horne

Dear Parent/Guardian,

I look forward to having a great year! I feel fortunate to have your learner in my class this semester and hope that you will contact me should you have any concerns. With your learner, please read the policies above that go with this form, then sign and date this signature page and return this form as soon as possible. Please provide a current email address and phone number at which I can contact you should the need arise. Please contact me with any concerns. It is going to be a GREAT semester!!!

Thank you,

Ashley Horne

My child and I have read and discussed the classroom syllabus.

Student Name (Print) _____ Date _____

Student Signature _____ Date _____

Parent/Guardian Name (Print) _____ Date _____

Parent/Guardian Signature _____ Date _____

Email Address(es) _____

Phone number(s) _____ (Cell) _____ (Home) _____ (Work)

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Course Description:

This course focuses on the composition, structure, and behavior of matter. Due to the emphasis on advanced problem-solving, incorporating algebraic skills, and a variety of laboratory activities, **it is highly recommended that the students have earned a "B" or higher average in Algebra 1**. This course meets the physical science graduation requirement.

Course Objectives:

Students will:

- Differentiate among pure substances, mixtures, elements, and compounds.
- Describe the structure of carbon chains, branched chains, and rings.
- Use the periodic table to identify periodic trends, including atomic radii, ionization energy, electronegativity, and energy levels.
- Describe solubility in terms of energy changes associated with the solution process.
- Use the kinetic theory to explain states of matter, phase changes, solubility, and chemical reactions.
- Solve stoichiometric problems involving relationships among number of particles, moles, and masses of reactants and products in a chemical reaction.
- Explain the behavior of ideal gasses in terms of pressure, volume, temperature, and number of particles.
- Distinguish among endothermic and exothermic physical and chemical changes.
- Distinguish between chemical and nuclear reactions.

Classroom Rules and Expectations:

1. **BE ON TIME**. Tardy means that you are not in the room and getting seated when the bell rings. *JCHS policy governs the consequences for tardiness*. In the event of an emergency hybrid schedule or virtual schedule as mandated by MCBOE, students will continue participating in class DURING the regular assigned time. Student access to assignment directions and Webex virtual meeting links will be provided by the teacher at the beginning of class.
2. **BE RESPECTFUL**: Practice courtesy and mutual respect. Treat others as you would like to be treated. The classroom and laboratory are to be regarded as a safe and supportive learning environment.
3. **BE PREPARED**: Mentally focused on reaching your goals and following class expectations; and physically bringing proper materials EVERY DAY. Cameras should remain on and microphone access should be readily available during live lessons to better facilitate class discussion. Any special circumstances must be communicated and approved by the instructor.
4. **BE RESOURCEFUL**: Thoroughly review assignments, videos, textbooks, and notes to answer questions before asking me.

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Accommodations:

Requests for accommodations for this course or any school event are welcomed from students and parents.

Concerning Laptop Utilization:

1. Student laptops should not be hard-wired to the network or have print capabilities. 2. Use of discs, flash drives, jump drives, or other USB devices will not be allowed on Madison City computers. 3. Neither the teacher, nor the school is responsible for broken, stolen, or lost laptops. 4. Laptops and other electronic devices will be used at the individual discretion of the teacher.

Grading Policy:

Test grades will account for 70% of the 9-weeks grade, with the remaining 30% being determined by quiz/daily grades. The grading scale is as follows: A (90-100%), B (80-89), C (70-79), D (65-69), and F (below 65). Grades will be a reflection of mastery of the standards. Make sure all absences are excused as class work can be made up and graded for excused absences only. The final exam counts for 20% of the final grade.

Make-Up Work Policy:

This course will follow the MCS Code of Conduct for make-up work. Make-up test time is once per week on a day determined by the instructor.

Course Materials:

- Scientific Calculator *(not a phone* many students like the Texas Instruments TI-30XIIS or the Casio FX-115ESPLUS
- Roll of Paper Towels *(for lab use)*
- Pencils
- 3 ring binder (at least one inch)

Texts/Required Readings:

Text: Introductory Chemistry: A Foundation, Zumdahl and DeCoste, 2015

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18 - WEEK PLAN*	
WEEK 1	Safety, Classification and Investigation of Matter
WEEK 2	Investigation of Matter/Units of Measure, Accuracy, Precision and Dimensional Analysis
WEEK 3	Investigation of Matter/Units of Measure, Accuracy, Precision and Dimensional Analysis
WEEK 4	Phase Changes and Heat
WEEK 5	KMT and Gases
WEEK 6	KMT and Gases
WEEK 7	Atomic History and Structure
WEEK 8	Chemical Bonding
WEEK 9	Chemical Bonding
WEEK 10	Nature of Electrons and Trends on the Periodic Table
WEEK 11	Nature of Electrons and Trends on the Periodic Table
WEEK 12	Chemical Reactions and Equations
WEEK 13	Chemical Reactions and Equations
WEEK 14	Stoichiometry
WEEK 15	Chemical Reactions and Equations
WEEK 16	Solutions
WEEK 17	Acids and Bases
WEEK 18	Equilibria/ Thermodynamics/Rates of Reaction

*** This syllabus serves as a guide for both the teacher and student; however, during the term it may become necessary to make additions, deletions or substitutions.**